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A chemist weighed out 5.14 g of a mixture containing unknown

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amounts of BaO(s) and CaO(s) and placed the sample in a 1.50-L flask containing $\text{CO}_2\text{(g)}$ at 30.08C and 750. torr. After the reaction to form $\text{BaCO}_3\text{(s)}$ and $\text{CaCO}_3\text{(s)}$ was completed, the pressure of $\text{CO}_2\text{(g)}$ remaining was 230. torr. Calculate the mass percentages of

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